



## An Introduction to the Study of Organic Chemistry

By Hans Thacher Clarke

Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1914 Excerpt: .of ferric salts, due to the formation of double salts such as  $9 \text{KSCN} \cdot 2 \text{Fe}(\text{SCN})_8$ . The thiocyanates are more stable than the cyanates, and remain without decomposition in aqueous solution for an indefinite period. Ammonium thiocyanate  $\text{NH}_4 \cdot \text{S} \cdot \text{C} \cdot \text{N}$ : N is formed on treating a solution of hydrogen cyanide with yellow ammonium sulphide; a more convenient method of preparation is to heat carbon bisulphide with alcoholic ammonia:  $\text{CS}_2 + 4\text{NH}_3 = \text{NH}_4 \cdot \text{S} \cdot \text{CN} + (\text{NH}_4)_2\text{S}$ . It is a crystalline substance, soluble in water and in alcohol, which melts at about 1500. On heating to 1700-1800 it undergoes an intramolecular rearrangement similar to that of ammonium cyanate, yielding thiocarbamide (p. 183). Alkyl thiocyanates  $\text{R} \cdot \text{S} \cdot \text{C} \cdot \text{N}$  in which an alkyl radicle replaces the metal, may be prepared by two different methods, which clearly show their constitution: firstly by the action of alkyl halides, or alkyl esters of inorganic salts, such as metallic...



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